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Bis(1*H*-benzimidazol-1-yl)methane monohydrate

Tao Shi,[a] Shouwen Jin,[b]* Jianlong Zhu,[b] YingJia Liu[b] and ChuanChuan Shi[b]

[a]Faculty of Science, ZheJiang A & F University, Lin'An 311300, People's Republic of China, and [b]Tianmu College of ZheJiang A & F University, Lin'An 311300, People's Republic of China
Correspondence e-mail: jinsw@zafu.edu.cn

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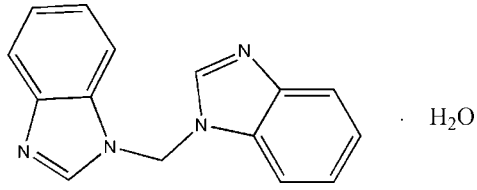
Key indicators: single-crystal X-ray study; T = 298 K; mean σ(C–C) = 0.006 Å; R factor = 0.057; wR factor = 0.181; data-to-parameter ratio = 13.0.

In the title compound, $C_{15}H_{12}N_4\cdot H_2O$, the organic molecule displays approximate non-crystallographic twofold symmetry: the dihedral angle between the benzimidazole ring systems is 81.37 (12)°. In the crystal, the components are linked by O—H···N hydrogen bonds, forming chains propagating in [101]. Aromatic π–π stacking [centroid–centroid separation = 3.595 (2) Å] helps to consolidate the structure.

Related literature

For background to coordination polymers containing bridged imidazole systems, see: Jin & Chen (2007); Ma *et al.* (2003). For the synthesis, see: Lavandera *et al.* (1988).



Experimental

Crystal data

$C_{15}H_{12}N_4\cdot H_2O$
M_r = 266.30
Triclinic, $P\overline{1}$
a = 7.3035 (6) Å
b = 8.9731 (8) Å
c = 11.1943 (10) Å
α = 103.578 (2)°
β = 103.408 (2)°
γ = 96.934 (1)°
V = 681.67 (10) Å^3
Z = 2
Mo $K\alpha$ radiation
μ = 0.09 mm^{-1}
T = 298 K
0.40 × 0.38 × 0.23 mm

Data collection

Bruker SMART CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2002)
T_{min} = 0.966, T_{max} = 0.980
3472 measured reflections
2350 independent reflections
1280 reflections with $I > 2\sigma(I)$
R_{int} = 0.026

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.057
$wR(F^2)$ = 0.181
S = 1.03
2350 reflections
181 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.20 e Å^{-3}
$\Delta\rho_{min}$ = −0.23 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O1—H1D···N4^{i}	0.85	2.14	2.940 (4)	157
O1—H1C···N2^{ii}	0.85	2.12	2.923 (3)	157

Symmetry codes: (i) $x, y, z-1$; (ii) $x+1, y, z$.

Data collection: *SMART* (Bruker, 2002); cell refinement: *SAINT* (Bruker, 2002); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB6425).

References

Bruker (2002). *SADABS*, *SMART* and *SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.
Jin, S. W. & Chen, W. Z. (2007). *Inorg. Chim. Acta*, **12**, 3756–3764.
Lavandera, J. L., Cabildo, P. & Claramunt, R. M. (1988). *J. Heterocycl. Chem.* **25**, 771–778.
Ma, J. F., Yang, J., Zheng, G. L., Li, L. & Liu, J. F. (2003). *Inorg. Chem.* **42**, 7531–7534.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.

supplementary materials

Acta Cryst. (2011). E**67**, o2943 [**doi:10.1107/S1600536811041572**]

Bis(1*H*-benzimidazol-1-yl)methane monohydrate

T. Shi, S. Jin, J. Zhu, Y. J. Liu and C. C. Shi

Comment

Bridged imidazole derivatives are a good choice of a N-donor ligand, and the flexible nature of the spacers allows the ligands to bend and rotate when coordinating to metal centers so as to conform to the coordination geometries of the metal ions. Significant progress has been achieved by us (Jin *et al.*, 2007) and others (Ma *et al.*, 2003) in this area.

As an extension of our research in bridged imidazole derivatives, here in this paper, we report the structure of the title compound, (I).

The r.m.s. deviations of the two benzimidazole rings are 0.003 Å and 0.007Å. They make dihedral angle of 81.37 (12)° with each other, indicating the almost perpendicular arrangement of both rings.

In the crystal, the water molecule and the bis(*N*-benzimidazolyl)methane molecule are connected together by the O—H···N hydrogen bonds to form a chain.

Experimental

The starting material bis(*N*-benzimidazolyl)methane was prepared according to the published procedure (Lavandera *et al.*, 1988). Crystals of bis(*N*-benzimidazolyl)methane monohydrate were formed during an experiment to recrystallize the title compound. A solid of bis(*N*-benzimidazolyl)methane (24.8 mg, 0.10 mmol) in 4 ml of dmf and 1 ml of water was stirred for about 1 h at room temperature to dissolve it, then the solution was filtered into a test tube. The solution was left standing at room temperature for three weaks, colorless block crystals were isolated after slow evaporation of the solution in air at ambient temperature. The crystals were collected and dried in air to give the title compound.

Refinement

H atoms bonded to the O atoms were located in a difference Fourier map, the O—H distance was kept 0.85 Å and refined isotropically. Other H atoms were positioned geometrically with C—H = 0.93–0.97 Å, and constrained to ride on their parent atoms with U_{iso}(H) = 1.2 U_{eq}(C).

Figures

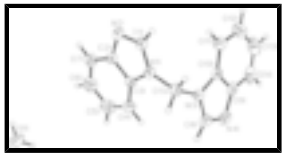

Fig. 1. The structure of the title compound, showing displacement ellipsoids drawn at the 30% probability level.

Fig. 2. Two-dimensional corrugated sheet structure formed through π-π interactions.

Bis(1*H*-benzimidazol-1-yl)methane monohydrate

Crystal data

$C_{15}H_{12}N_4 \cdot H_2O$	$Z = 2$
$M_r = 266.30$	$F(000) = 280$
Triclinic, $P\overline{1}$	$D_x = 1.297$ Mg m^{-3}
Hall symbol: -P 1	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
$a = 7.3035$ (6) Å	Cell parameters from 733 reflections
$b = 8.9731$ (8) Å	$\theta = 2.4$–$21.4°$
$c = 11.1943$ (10) Å	$\mu = 0.09$ mm^{-1}
$\alpha = 103.578$ (2)°	$T = 298$ K
$\beta = 103.408$ (2)°	Block, colorless
$\gamma = 96.934$ (1)°	0.40 × 0.38 × 0.23 mm
$V = 681.67$ (10) Å^3	

Data collection

Bruker SMART CCD diffractometer	2350 independent reflections
Radiation source: fine-focus sealed tube	1280 reflections with $I > 2\sigma(I)$
graphite	$R_{int} = 0.026$
φ and ω scans	$\theta_{max} = 25.0°$, $\theta_{min} = 2.4°$
Absorption correction: multi-scan (*SADABS*; Bruker, 2002)	$h = -7 \rightarrow 8$
$T_{min} = 0.966$, $T_{max} = 0.980$	$k = -9 \rightarrow 10$
3472 measured reflections	$l = -13 \rightarrow 12$

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.057$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.181$	H-atom parameters constrained
$S = 1.03$	$w = 1/[\sigma^2(F_o^2) + (0.0662P)^2 + 0.2485P]$ where $P = (F_o^2 + 2F_c^2)/3$
2350 reflections	$(\Delta/\sigma)_{max} < 0.001$
181 parameters	$\Delta\rho_{max} = 0.20$ e Å^{-3}
0 restraints	$\Delta\rho_{min} = -0.23$ e Å^{-3}

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
N1	−0.0524 (4)	0.0854 (3)	0.6722 (2)	0.0493 (7)
N2	−0.2207 (4)	0.1766 (3)	0.5192 (2)	0.0608 (8)
N3	0.0643 (4)	0.1058 (3)	0.8964 (2)	0.0519 (7)
N4	0.2853 (4)	0.2315 (4)	1.0804 (3)	0.0759 (9)
O1	0.5456 (4)	0.1029 (3)	0.2553 (2)	0.0940 (9)
H1C	0.6008	0.1485	0.3334	0.113*
H1D	0.4776	0.1620	0.2238	0.113*
C1	−0.2270 (5)	0.0924 (4)	0.6003 (3)	0.0593 (9)
H1	−0.3411	0.0421	0.6079	0.071*
C2	−0.0259 (5)	0.2298 (4)	0.5402 (3)	0.0488 (8)
C3	0.0822 (4)	0.1746 (3)	0.6359 (3)	0.0445 (7)
C4	0.2790 (5)	0.2106 (4)	0.6748 (3)	0.0642 (10)
H4	0.3493	0.1728	0.7380	0.077*
C5	0.3680 (6)	0.3060 (5)	0.6156 (4)	0.0838 (12)
H5	0.5008	0.3343	0.6404	0.101*
C6	0.2613 (6)	0.3602 (5)	0.5194 (4)	0.0806 (12)
H6	0.3252	0.4224	0.4805	0.097*
C7	0.0655 (6)	0.3244 (4)	0.4807 (3)	0.0637 (10)
H7	−0.0041	0.3617	0.4170	0.076*
C8	0.2462 (5)	0.1256 (5)	0.9711 (3)	0.0686 (10)
H8	0.3356	0.0683	0.9463	0.082*
C9	0.1143 (5)	0.2873 (4)	1.0789 (3)	0.0583 (9)
C10	−0.0250 (4)	0.2100 (4)	0.9655 (3)	0.0482 (8)
C11	−0.2091 (5)	0.2400 (4)	0.9394 (3)	0.0622 (9)
H11	−0.3005	0.1871	0.8634	0.075*
C12	−0.2505 (6)	0.3520 (5)	1.0315 (4)	0.0808 (12)
H12	−0.3732	0.3750	1.0175	0.097*
C13	−0.1141 (8)	0.4315 (5)	1.1446 (4)	0.0867 (13)
H13	−0.1473	0.5066	1.2045	0.104*
C14	0.0688 (7)	0.4015 (5)	1.1699 (3)	0.0788 (12)
H14	0.1599	0.4558	1.2456	0.095*
C15	−0.0163 (5)	0.0014 (4)	0.7692 (3)	0.0576 (9)
H15A	−0.1356	−0.0618	0.7664	0.069*

H15B	0.0713	−0.0682	0.7505	0.069*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
N1	0.0491 (16)	0.0624 (17)	0.0391 (14)	0.0136 (13)	0.0125 (12)	0.0168 (13)
N2	0.0571 (19)	0.087 (2)	0.0422 (15)	0.0220 (15)	0.0100 (13)	0.0243 (15)
N3	0.0533 (17)	0.0689 (18)	0.0401 (14)	0.0216 (14)	0.0122 (13)	0.0225 (13)
N4	0.065 (2)	0.112 (3)	0.0501 (18)	0.0118 (18)	0.0035 (15)	0.0338 (19)
O1	0.103 (2)	0.098 (2)	0.0688 (16)	0.0330 (17)	−0.0108 (15)	0.0273 (15)
C1	0.049 (2)	0.085 (3)	0.0428 (18)	0.0131 (18)	0.0125 (16)	0.0165 (18)
C2	0.056 (2)	0.054 (2)	0.0377 (16)	0.0181 (16)	0.0131 (15)	0.0107 (15)
C3	0.0490 (19)	0.0477 (18)	0.0417 (16)	0.0188 (15)	0.0159 (14)	0.0132 (14)
C4	0.050 (2)	0.079 (3)	0.071 (2)	0.0227 (18)	0.0155 (18)	0.031 (2)
C5	0.057 (2)	0.092 (3)	0.110 (3)	0.011 (2)	0.031 (2)	0.036 (3)
C6	0.087 (3)	0.083 (3)	0.094 (3)	0.020 (2)	0.045 (3)	0.044 (2)
C7	0.081 (3)	0.065 (2)	0.057 (2)	0.025 (2)	0.026 (2)	0.0269 (18)
C8	0.056 (2)	0.103 (3)	0.062 (2)	0.033 (2)	0.0165 (19)	0.043 (2)
C9	0.066 (2)	0.069 (2)	0.0407 (18)	0.0001 (19)	0.0132 (17)	0.0239 (18)
C10	0.053 (2)	0.058 (2)	0.0396 (17)	0.0110 (16)	0.0147 (15)	0.0212 (15)
C11	0.061 (2)	0.070 (2)	0.057 (2)	0.0191 (19)	0.0156 (17)	0.0177 (18)
C12	0.083 (3)	0.079 (3)	0.096 (3)	0.029 (2)	0.046 (3)	0.025 (3)
C13	0.117 (4)	0.067 (3)	0.080 (3)	0.010 (3)	0.050 (3)	0.009 (2)
C14	0.106 (4)	0.078 (3)	0.043 (2)	−0.012 (2)	0.021 (2)	0.011 (2)
C15	0.074 (2)	0.063 (2)	0.0438 (18)	0.0218 (18)	0.0201 (16)	0.0208 (17)

Geometric parameters (Å, °)

N1—C1	1.358 (4)	C5—C6	1.397 (5)
N1—C3	1.385 (4)	C5—H5	0.9300
N1—C15	1.455 (4)	C6—C7	1.368 (5)
N2—C1	1.316 (4)	C6—H6	0.9300
N2—C2	1.391 (4)	C7—H7	0.9300
N3—C8	1.362 (4)	C8—H8	0.9300
N3—C10	1.393 (4)	C9—C10	1.395 (4)
N3—C15	1.450 (4)	C9—C14	1.396 (5)
N4—C8	1.308 (4)	C10—C11	1.381 (4)
N4—C9	1.398 (4)	C11—C12	1.377 (5)
O1—H1C	0.8500	C11—H11	0.9300
O1—H1D	0.8500	C12—C13	1.387 (6)
C1—H1	0.9300	C12—H12	0.9300
C2—C7	1.394 (4)	C13—C14	1.373 (6)
C2—C3	1.403 (4)	C13—H13	0.9300
C3—C4	1.376 (4)	C14—H14	0.9300
C4—C5	1.387 (5)	C15—H15A	0.9700
C4—H4	0.9300	C15—H15B	0.9700
C1—N1—C3	106.7 (2)	C2—C7—H7	121.2
C1—N1—C15	126.0 (3)	N4—C8—N3	114.7 (3)

C3—N1—C15	127.3 (3)	N4—C8—H8	122.7
C1—N2—C2	103.8 (3)	N3—C8—H8	122.7
C8—N3—C10	105.8 (3)	C10—C9—C14	119.3 (3)
C8—N3—C15	126.7 (3)	C10—C9—N4	110.3 (3)
C10—N3—C15	127.4 (3)	C14—C9—N4	130.3 (3)
C8—N4—C9	103.9 (3)	C11—C10—N3	132.0 (3)
H1C—O1—H1D	108.9	C11—C10—C9	122.7 (3)
N2—C1—N1	114.2 (3)	N3—C10—C9	105.3 (3)
N2—C1—H1	122.9	C12—C11—C10	116.6 (3)
N1—C1—H1	122.9	C12—C11—H11	121.7
N2—C2—C7	129.1 (3)	C10—C11—H11	121.7
N2—C2—C3	110.7 (3)	C11—C12—C13	121.8 (4)
C7—C2—C3	120.1 (3)	C11—C12—H12	119.1
C4—C3—N1	133.1 (3)	C13—C12—H12	119.1
C4—C3—C2	122.2 (3)	C14—C13—C12	121.3 (4)
N1—C3—C2	104.7 (3)	C14—C13—H13	119.4
C3—C4—C5	117.0 (3)	C12—C13—H13	119.4
C3—C4—H4	121.5	C13—C14—C9	118.2 (4)
C5—C4—H4	121.5	C13—C14—H14	120.9
C4—C5—C6	121.1 (4)	C9—C14—H14	120.9
C4—C5—H5	119.5	N3—C15—N1	112.2 (3)
C6—C5—H5	119.5	N3—C15—H15A	109.2
C7—C6—C5	121.9 (4)	N1—C15—H15A	109.2
C7—C6—H6	119.1	N3—C15—H15B	109.2
C5—C6—H6	119.1	N1—C15—H15B	109.2
C6—C7—C2	117.6 (3)	H15A—C15—H15B	107.9
C6—C7—H7	121.2		
C2—N2—C1—N1	−0.4 (3)	C15—N3—C8—N4	177.6 (3)
C3—N1—C1—N2	0.6 (3)	C8—N4—C9—C10	0.2 (4)
C15—N1—C1—N2	179.9 (3)	C8—N4—C9—C14	179.9 (3)
C1—N2—C2—C7	179.9 (3)	C8—N3—C10—C11	−178.7 (3)
C1—N2—C2—C3	0.0 (3)	C15—N3—C10—C11	3.4 (5)
C1—N1—C3—C4	179.8 (3)	C8—N3—C10—C9	0.5 (3)
C15—N1—C3—C4	0.5 (5)	C15—N3—C10—C9	−177.5 (3)
C1—N1—C3—C2	−0.5 (3)	C14—C9—C10—C11	−0.9 (5)
C15—N1—C3—C2	−179.8 (3)	N4—C9—C10—C11	178.8 (3)
N2—C2—C3—C4	−180.0 (3)	C14—C9—C10—N3	179.8 (3)
C7—C2—C3—C4	0.2 (4)	N4—C9—C10—N3	−0.4 (3)
N2—C2—C3—N1	0.3 (3)	N3—C10—C11—C12	179.3 (3)
C7—C2—C3—N1	−179.6 (3)	C9—C10—C11—C12	0.2 (5)
N1—C3—C4—C5	−179.9 (3)	C10—C11—C12—C13	0.3 (5)
C2—C3—C4—C5	0.4 (5)	C11—C12—C13—C14	−0.2 (6)
C3—C4—C5—C6	−1.0 (5)	C12—C13—C14—C9	−0.5 (6)
C4—C5—C6—C7	1.1 (6)	C10—C9—C14—C13	1.0 (5)
C5—C6—C7—C2	−0.5 (6)	N4—C9—C14—C13	−178.7 (3)
N2—C2—C7—C6	−180.0 (3)	C8—N3—C15—N1	−110.3 (3)
C3—C2—C7—C6	−0.1 (5)	C10—N3—C15—N1	67.3 (4)
C9—N4—C8—N3	0.2 (4)	C1—N1—C15—N3	−112.9 (3)
C10—N3—C8—N4	−0.4 (4)	C3—N1—C15—N3	66.3 (4)

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O1—H1D···N4^{i}	0.85	2.14	2.940 (4)	157
O1—H1C···N2^{ii}	0.85	2.12	2.923 (3)	157

Symmetry codes: (i) *x*, *y*, *z*−1; (ii) *x*+1, *y*, *z*.

Fig. 1

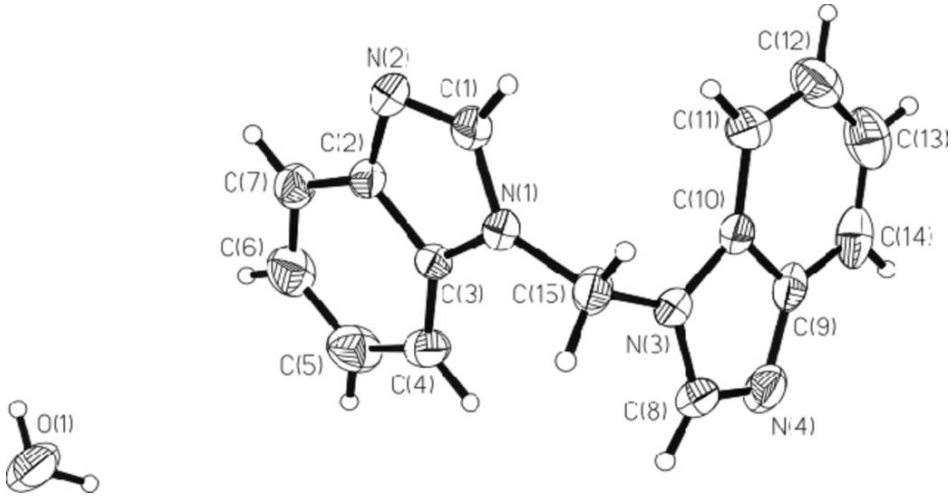

Fig. 2